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PCT/EP 02/11963

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Blatt 2 der Bescheinigung  
Sheet 2 of the certificate  
Page 2 de l'attestation



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The registration of the changes has taken effect from 31.01.2003.

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## SOLID OXIDE FUEL CELL

The present invention relates to a solid fuel cell, to a cermet material, and to the process for the preparation of said cermet.

5 Solid-oxide fuel cells (SOFCs) convert chemical energy into electrical energy with high efficiency and low emission of pollutants. Although the introduction of a "green energy" might seem an attractive scenario, its implementation is beset with technical and economic difficulties.

10 The most common anodes materials for solid oxide fuel cells comprise nickel (Ni) cermets (ceramic and metallic composite materials) prepared by high-temperature calcination of NiO and ceramic powders, usually yttria-stabilized zirconia (YSZ) powders. These Ni-cermets perform with H<sub>2</sub> fuels and allow internal steam reforming of hydrocarbons if there is sufficient water in the feed to the anode. Because Ni catalyzes the formation of graphite fibers in dry methane, it is necessary to operate anodes at steam/methane ratios greater than 3, as from WO 00/52780 (In the name of Gas Research Institute).

15 S.J.A. Livermore et al. *Journal of Power Sources*, vol. 86 (2000), 411-416, refers to a cermet anode for SOFC made of nickel and ceria-gadolinia (CGO). This anode performs at 600°C using 10% H<sub>2</sub>/N<sub>2</sub> as the fuel.

20 A. Müller et al., Proc. of the 3rd European Solid Fuel Cell Forum, Nantes France, June 1998, 353-362, relate to a Ni-YSZ anode fuel cell. It is envisaged a degradation of the anode related to microstructural changes occurring during operation. The nickel particles have a mean diameter of about 0.5 µm, and are homogeneously distributed in the anode. After long term operation at high current density and fuel utilization (H<sub>2</sub>+H<sub>2</sub>O), the agglomeration of the nickel particles leads to a decrease of the amount of three-phase boundary (TPB), resulting in an increase in the anode losses.

25 A.C. Müller et al., HTMC IUPAC Jülich 2000 suggest that the degradation described by the previous document could be prevented by a multilayer anode whose divers layers differ in their microstructure to fulfill the locally different requirements for SOFC anodes. In particular, the content of Ni and the Ni particle size should increase from first layer (that in contact with the elec-

30

trolyte) to last layer, thus increasing electronic conductivity, TEC (Thermal Expansion Coefficient) and porosity. The YSZ content should simultaneously decrease. The cermet samples were prepared by mixing 65-85 mol% NIO powder with YSZ powder and sintering them in air at 1300°C for 5 hours. The particle size of the metallic portion was 0.5-8  $\mu\text{m}$ .

The use of nickel as the metallic component of a cermet anode is advantageous, but its performance drops in short time, especially when a dry hydrocarbon is the fuel, due to graphite formation.

R.J. Gorte et al., *Adv. Mater.*, 2000, vol.12, No. 19, 1465-1469, propose to substitute nickel with copper (Cu) in a cermet wherein the ceramic portion is YSZ. Other components, including ceria ( $\text{CeO}_2$ ), can be added to the metallic portion. In this configuration the role of  $\text{CeO}_2$  is mainly to provide catalytic activity for the oxidation of hydrocarbons. As shown in Figure 4a of this paper, the cell prepared with Cu but without ceria exhibits poor performance at 700°C, especially when methane is used as fuel.

C. Lu et al., High Temperature Materials, Proceedings volume 2002-5, Ed. S.C. Singhal, Pacific Northwest National Laboratory, Richland, Washington, USA, relate to a Cu-SDC (samaria-doped ceria) anode composite cell performing  $\text{H}_2$  and butene fuels at 700°C.

From the above studies, it results that copper alone cannot be an efficient substitute for nickel as its performance is insufficient, in particular with fuels such as dry hydrocarbons.

H. Kim et al., *J. Electrochem. Soc.*, vol.149 (3), A247-A250 (2002) examine the use of Cu-Ni alloys as anode component for the direct oxidation of methane in SOFC at about 800°C. The ceramic portion, which in this case is YSZ, is made by casting a tape with graphite pore formers over a green tape of YSZ without pore formers, firing the two-layered tape to about 1500°C. The porous anode layer was then impregnated with an aqueous solution of  $\text{Ce}(\text{NO}_3)_3 \cdot \text{H}_2\text{O}$  and calcinated at about 500°C to decompose the nitrate ions and form  $\text{CeO}_2$ . After the addition of ceria, the porous layer was impregnated with a mixed, aqueous solution of  $\text{Cu}(\text{NO}_3)_2 \cdot \text{H}_2\text{O}$  and  $\text{Ni}(\text{NO}_3)_2 \cdot \text{H}_2\text{O}$  having the

desired Cu:Ni ratio. Finally the wafer was again heated to about 500°C in air to decompose the nitrates and reduced in flowing H<sub>2</sub> at about 900°C.

5 The Applicant has faced the problem of providing a solid oxide fuel cell which is able to show high efficiency and to maintain its performance over time, particularly in terms of a low overpotential in a wide range of temperatures. Moreover, the fuel cell should be able to show the above characteristics when fed with different fuels. Endurance of performance is particularly important when a dry hydrocarbon is used as fuel, since it tends to form graphite fibers on the metallic portion of the cermet anode, which eventually annihilate the  
10 fuel cell activity.

The Applicant has now found that by using a nickel-copper alloy as metallic portion of a cermet anode, and by reducing the average size of the particles constituting said alloy, the resulting SOFC shows enduring efficiency when fuelled with different fuels, including hydrogen and hydrocarbons, in a  
15 wide range of operating temperatures, and particularly at low temperatures, namely from 500°C to 700°C. Particularly, when a dry hydrocarbon is used as fuel, deposition of graphite fibers is remarkably reduced. This result is particularly surprising since a reduction of the average particle size in the metallic component of the cermet would have been expected to increase the  
20 catalytic activity also with respect to those side reactions which cause formation of the graphite fibers.

The present invention thus relates to a solid oxide fuel cell including

- a cathode;
- at least an electrolyte membrane, and
- 25 - an anode comprising a ceramic material and a Cu/Ni alloy, said Cu/Ni alloy having an average particle size not higher than 20 nm.

Preferably the anode of the invention comprises a copper/nickel alloy wherein said alloy has an average particle size not higher than 16 nm. More preferably said average particle size is not lower than 1 nm.

The Cu/Ni alloy of the anode of the invention can show a mean surface area higher than 20 m<sup>2</sup>/g, preferably higher than 30 m<sup>2</sup>/g, and more preferably higher than 40 m<sup>2</sup>/g.

5 Preferably in the anode of the invention the Cu/Ni alloy has a copper content of from about 1% by weight to about 99% by weight, more preferably, and even more preferably from about 40% by weight to about 60% by weight.

10 Preferably in the anode of the present invention the Cu/Ni alloy has a nickel content of from about 1% by weight to about 99% by weight, more preferably from about 30% by weight to about 70% by weight, and even more preferably of about 50% by weight.

Said Cu/Ni alloy can comprise an additional metal, for example, cobalt, manganese, molybdenum or aluminum.

15 As additional metal is herein intended an element belonging to one of the classes from 3 to 13 of the periodic table of elements according to Chemical and Engineering News, 63(5), 27, 1985, lanthanides series included.

20 The ceramic material of the anode of the invention can be selected from yttria-stabilized zirconia (YSZ), cerium gadolinium oxide (CGO), samarium-doped ceria (SDC), mixed lanthanum and gallium oxides. Preferably the ceramic material is cerium gadolinium oxide (CGO).

The ceramic material of the anode of the present invention can show a particle size not higher than 50 nm, preferably from about 1 to about 25 nm.

25 Optionally, said ceramic material is doped with at least one cation selected from calcium, magnesium, strontium, lanthanum, yttrium, ytterbium, neodymium and dysprosium.

Optionally, the Cu/Ni alloy of the invention comprises cerium oxide (CeO<sub>2</sub>), optionally added with additives like cobalt.

30 Preferably, the solid oxide fuel cell according to the invention performs by feeding the anode with a fuel selected from hydrogen; an alcohol such as methanol, ethanol, propanol; a hydrocarbon such as methane, ethane, butene; carbon dioxide, carbon monoxide, natural gas, reformed natural gas,

biogas, syngas and mixture thereof, either in the presence of water or substantially dry. Preferred by the present invention is substantially dry methane.

The solid oxide fuel cell according to the present invention can be operated in a wide range of temperatures, usually ranging from 450°C to 800°C, and preferably from 500°C to 700°C

In another aspect the present invention relates to a cermet comprising a Cu/Ni alloy having a particle size not higher than 20 nm.

Both the metallic and the ceramic portion of the cermet anode of the present invention can be prepared from the corresponding metal salts, which may be compounded in a solid solution of the oxides thereof.

In a further aspect the present invention relates to a process for preparing a cermet including a ceramic material and a metallic material comprising a Cu/Ni alloy, said process comprising the steps of:

- a) producing a precursor of the metallic material;
- b) producing the ceramic material;
- c) combining said precursor and ceramic material to obtain a composite;
- d) reducing said composite

wherein step a) comprises the phases of

- a-1) dissolving a hydrosoluble salt of Ni and a hydrosoluble salt of Cu in water;
- a-2) adding a chelating agent to the solution resulting from step a-1);
- a-3) adding an oxidizing agent to the solution resulting from step a-2);
- a-4) isolating said precursor.

Optionally, the pH of the solution resulting from phase a-2) is adjusted at a value higher than about 5.

Preferably, phase d) is carried out by hydrogen at a temperature ranging between about 400°C and about 1000°C, more preferably between about 500°C and about 800°C.



Step b) can include phases analogous to those from a-1) to a-4) of the process above, starting from the hydrosoluble salts corresponding to the oxide of the desired ceramic material. Phase a-4) can be followed by a crystallization step, for example at a temperature ranging between about 200°C and about 700°C, more preferably between about 300°C and about 500°C.

Preferably, the oxides of the present invention are prepared starting from hydrosoluble salts of the desired metals which are dissolved in water and added with a chelating agent, for example, EDTA, oxalic, citric, acetic acid and the organic salts thereof, while maintaining the reaction mixture at a suitable pH, for example, higher than about 5.

Oxidation is then carried out, e.g. by addition of a peroxide, such as hydrogen peroxide, and co-precipitate of amorphous metal oxides is obtained.

This co-precipitate comprises very fine amorphous particles substantially free from any crystallographic ordering as revealed by XRD (X-ray diffraction) and TEM (transmission electron microscopy) analyses, as shown in the examples. The morphological and structural features obtained through this preferred method allow a superior extension of the three-phase boundary (TPB) zone, advantageous to the performance of the SOFC.

After thermal treatment of the precursor, in air or inert atmosphere, for example helium, a solid solution of the metallic oxides intimately admixed on an atomic scale, with fine particles size, is obtained. The particle size can range from about 3 to about 20 nm, preferably from about 4 to about 7 nm, more preferably of about 5 nm.

In the case the ceramic portion of the cermet anode according to the invention is prepared through the above mentioned process, crystallization of the amorphous oxide precursor, for example at a temperature ranging between about 200°C and about 700°C, more preferably between about 300°C and about 500°C, can yield a ceramic with small particle size, for example ranging between about 6 and about 2nm.

The preparation of the cermet anode, i.e. the material system comprising a metallic and a ceramic phase, can be carried out as follows. Amorphous  $\text{NiCuO}_x$  oxide precursor, obtained as said above, and a ceramic powder, preferably CGO or SDC, are admixed, and a slurry is prepared by dispersing the reactants in an organic solvent, for example isopropanol, and further treated with ultrasounds. The mixture is heated for solvent evaporation, and a reduction, for example in  $\text{H}_2$  atmosphere, is carried out while heating, for example at a temperature ranging between about  $400^\circ\text{C}$  and about  $1000^\circ\text{C}$ , more preferably between about  $500^\circ\text{C}$  and about  $800^\circ\text{C}$ .

A solid oxide fuel cell of the invention can be prepared by applying said slurry of composite on an electrolyte membrane comprising a ceramic material, for example, CGO, SDC or YSZ.

A cathode for the solid oxide fuel cell of the invention can comprise a perovskite such as  $\text{La}_{1-x}\text{Sr}_x\text{MnO}_{3-\delta}$ , for example can be a  $\text{La}_{0.6}\text{Sr}_{0.4}\text{MnO}_3/\text{CGO}$ .

The invention will be further illustrated hereinafter with reference to the following examples and figures, wherein

- Figure 1 illustrates a schematic view of a solid oxide fuel cell;
- Figure 2 illustrates XRD patterns of (a) amorphous oxide precursor  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}$ , (b) crystalline oxide precursor  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}$  and (c)  $\text{Ni}_{0.58}\text{Cu}_{0.42}$  bulk alloy;
- Figure 3 shows temperature reduction profiles with  $\text{H}_2$  for (a)  $\text{CuO}$ , (b)  $\text{NiO}$  and (c)  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}$  crystalline precursor;
- Figure 4 illustrates catalytic activity vs temperature profile for methane steam reforming experiments carried out for  $\text{Ni}_{0.58}\text{Cu}_{0.42}$  bulk alloy and  $\text{Ni,Cu}$  metallic mixture;
- Figure 5 shows XRD pattern of a  $\text{Ni}_{0.58}\text{Cu}_{0.42}$  (\*)-CGO (o) cermet;
- Figure 6 shows XRD patterns of (a) crystalline  $\text{Ni}_{0.5}\text{Cu}_{0.5}\text{O}$ -CGO composite after thermal treatment at  $500^\circ\text{C}$  and (b)  $\text{Ni}_{0.5}\text{Cu}_{0.5}$  (\*)-CGO (o) cermet after reduction at  $900^\circ\text{C}$ ; (#) corresponds to  $\text{CuO}$  phase;

- Figure 7 shows electrochemical polarisation curves for  $\text{Ni}_{0.58}\text{Cu}_{0.42}$ -CGO cermet anode/CGO electrolyte interface in dry  $\text{H}_2$  and dry  $\text{CH}_4$  in the temperature range between  $650^\circ$  and  $800^\circ\text{C}$ ;
- Figure 8 illustrates IR free electrochemical polarisation curves for  $\text{Ni}_{0.58}\text{Cu}_{0.42}$ -CGO cermet anode/CGO electrolyte interface in dry  $\text{H}_2$  and dry  $\text{CH}_4$  in the temperature range between  $650^\circ$  and  $800^\circ\text{C}$ ;
- Figures 9 illustrates a SEM micrograph of a  $\text{Ni}_{0.58}\text{Cu}_{0.42}$ -CGO anodic cermet layer of the invention, in cross-section;
- Figures 10a and 10b show TEM images of, respectively, amorphous and crystalline  $\text{Ni}_{0.58}\text{Cu}_{0.42}$ -CGO cermet according to the invention;
- Figure 11 show a low resolution TEM micrograph of  $\text{Ni}_{0.58}\text{Cu}_{0.42}$ -CGO anodic cermet layer, according to the invention, after 25 h at  $250 \text{ mA cm}^{-2}$  in dry methane.

Figure 1 schematically illustrates a solid oxide fuel cell comprising anode 1, cathode 2, and electrolyte membrane 3. the arrows indicate the electron flow from anode to cathode during operation.

#### Example 1

##### Oxide precursor and alloy preparation

- a) NiCu alloys were prepared from reagent graded  $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  and  $\text{Cu}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  (Aldrich 99.99). Stoichiometric amounts of the metal nitrates (2.86 g of  $\text{Cu}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , 3.1 g of  $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ ) were dissolved in distilled water (50 ml) and then complexed at  $60^\circ\text{C}$  with an aqueous solution of oxalic acid (9.5 g in 200ml; Aldrich 99.99) at  $\text{pH}=6.5$  adjusted with  $\text{NaOH}$  0.1N. The molar ratio between complexing agent and the sum of the metal ions was 10. The complex formation was monitored by UV spectroscopy. The solution was heated to  $80^\circ\text{C}$ , and oxygen peroxide (400 ml, 20%, Carlo Erba) was then dropwise added until complete formation of a precipitate. The precipitate was filtrated, washed with distilled water, and dried at  $120^\circ\text{C}$  K for 12 hours.

- b) The powder obtained at step a) was then calcinated at 500°C in air for 0.5 hour, to yield a crystalline phase.
- c) The crystalline phase of step b) was heated at 500°C under H<sub>2</sub> atmosphere (H<sub>2</sub> for 30 min, 50 cc min<sup>-1</sup> g<sup>-1</sup> catalyst), in order to reduce the oxide phase to a metallic phase.

In order to verify the phase composition and structure at steps a), b) and c), the powders were analyzed by X-ray fluorescence (XRF) and X-ray diffraction (XRD) analyses.

XRF analysis (for composition) was carried out by Explorer Spectrometer (Bruker AXS, Germany) equipped with a Rh X-ray source. The instrument was equipped with 0.12° divergence collimator, LiF220 crystal analyzer and scintillation as well as proportional detectors.

XRD analysis (for structure and particle size determination) was carried out under Bragg-Brentano configuration with CuKalpha radiation with Xpert Diffractometer (Philips). The instrument was equipped with graphite monochromator. The analysis range was selected from 5° to 100° 2θ, the sweep rate was 1.5 ° min<sup>-1</sup>. Results are shown in Figure 2.

For each step product, the average particle size was calculated from XRD line broadening measurements using Scherrer equation. The results are set forth in Table 1, wherein the metal surface area was calculated from the particle size by the following formula:

$$MSA (m^2g^{-1}) = 6 \cdot 10^4 / (\rho \cdot d);$$

wherein  $\rho = (g \text{ cm}^{-3})$  is density and  $d (\text{\AA})$  particle size.

The formation of the metal solid solution was checked by XRD, monitoring the shifts on the diffraction peak assigned to the various reflections in particular the Ni (200) (JCPDS - ICCD data file, card n° 4-8509) and Cu (200) (JCPDS - ICCD data file, card n° 4-836) planes, and calculating the variation in the lattice parameter of the metallic alloy structure from Vegard law.

Table 1

Lattice parameter (a), particle size (d) and surface area (SA) of NiCuO<sub>x</sub> and NiCu phases

Phase / preparation step (ex. 1)	(Ni:Cu)	Lattice parameter $a_{fcc}/\text{\AA}$	Particle size $d/\text{nm}$	Surface area $SA/\text{m}^2 \text{g}^{-1}$
NiCuO <sub>x</sub> amorphous (step a)	90:10	-	< 1.5	-
	65:35	-	< 1.5	-
	58:42	-	< 1.5	-
NiCuO <sub>x</sub> crystalline (step b)	90:10	4.162	6.7	100.6
	65:35	4.178	4.1	164.4
	58:42	4.183	4.3	156.8
NiCu alloy (step c)	90:10	3.521	15.9	42.4
	65:35	3.549	13.4	50.3
	58:42	3.552	19.2	36.0

XRD analysis of the powder obtained from step a) did not show any important crystallographic reflection but only an amorphous scattering, as from Figure 2,a).

5 XRD analysis of the powder obtained from step b) showed the characteristic peaks of the faced centered cubic (fcc) structure of NiO shifted linearly as a function of the composition of the solid solution, as from Figure 2,b). No evidence of monocline CuO phase was observed indicating that Cu atoms occupy the same crystallographic positions of Ni in the fcc structure with a random distribution. No evidence of superlattice lines was detected.

10 XRD analysis of the powder obtained from step c) showed for the various compositions only the typical diffraction peaks of the fcc structure, as from Figure 2,c). The lattice parameter varied linearly between pure Ni and Cu metals as a function of relative composition.

15 Summarizing, at step a) an amorphous oxide with very small particle size (<1.5 nm) is formed (see Figure 2,a), at step b) a crystalline oxide solid solution is formed with a particle size around 5 nm (see Figure 2,b), and at step c) (see Figure 2,c) a single metallic phase is obtained (particle size 10-20 nm).

20 Example 2

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Temperature Programmed Reduction (TPR) on NiO, CuO and NiCuO<sub>x</sub> oxides

5 TPR experiments were carried out in a tubular quartz microreactor. A temperature sweep rate of 10 °C min<sup>-1</sup> was selected, the catalyst weight was 5 mg and a TCD detector was used to determine H<sub>2</sub> consumption. A stream of 5% H<sub>2</sub> in Ar at a flow rate of 30 ml/min was fed to the reactor.

10 Three different samples, CuO (Aldrich), NiO (Aldrich) and Ni<sub>0.58</sub>Cu<sub>0.42</sub>O<sub>x</sub> were subjected to TPR experiments with H<sub>2</sub>, in order to compare their reduction kinetics and confirm the effectiveness of the process of the invention in preparing a metallic alloy.

15 The recorded profiles of Figure 3 show that the onset for hydrogen consumption and the respective peak maximum did occur on the mixed phases at a much lower temperatures with respect to the single phases: 470 K (about 197°C) for Ni<sub>0.58</sub>Cu<sub>0.42</sub>O<sub>x</sub>, 503 K (about 230°C) for CuO and 626 K (about 353°C) for NiO, respectively. Further, the reduction peak for the mixed Ni<sub>0.58</sub>Cu<sub>0.42</sub>O<sub>x</sub> phases is much more narrow, pointing for a faster reduction kinetics, thus a higher affinity for hydrogen, effective for reforming.

Example 3

Catalytic activity for methane oxidation (reforming)

20 Methane reforming experiments were carried out in a packed bed tubular reactor (in-house made), where 15 mg of alloy catalyst of example 1,c) were diluted in 50 mg of quartz powder (Carlo Erba). Reactant feed was H<sub>2</sub>O/CH<sub>4</sub> (molar ratio 4:1), with a space velocity of 10<sup>5</sup> h<sup>-1</sup>.

25 Methane steam reforming experiments were carried out on a Ni<sub>0.58</sub>Cu<sub>0.42</sub> bulk alloy example 1,c) and on a Ni,Cu metallic mixture of same compositions prepared by grinding and ultrasonically mixing NiO and CuO oxides (Aldrich) followed by reduction under same conditions of Example 1,c).

30 Figure 4 shows that the onset temperature for the reaction on the Ni<sub>0.58</sub>Cu<sub>0.42</sub> alloy is significantly lower than for that on the Ni,Cu metallic mixture. The inflection point in the curve of catalytic activity vs. temperature for Ni<sub>0.58</sub>Cu<sub>0.42</sub> alloy is about 50 K (200°C) lower than that of Ni,Cu metallic mix-

ture and 100 K lower than that reported in the literature for supported Ni catalysts (see C.T Au, H.Y. Wang, H.C. Wan, J. Catalysis 158 (1996) 343).

Elemental analysis was used to determine the carbon content in the catalyst after catalytic and electrochemical experiments. Analyses were carried out with a Carlo Erba CHNSO elemental analyser. No evidence of carbon deposition was found after reforming experiments on the  $\text{Ni}_{0.58}\text{Cu}_{0.42}$  alloy.

#### Example 4

##### $\text{Ni}_{0.58}\text{Cu}_{0.42}$ -CGO (50:50 w/w) cermet preparation

$\text{Ce}_{0.9}\text{GdO}_{1.95}$  (CGO) (prepared by co-precipitation of the cerium and gadolinium nitrates with oxalic acid at pH=6, followed by thermal decomposition at 973K, as from Herle J.V., Horita T., Kawada T., Sato N., Yokokawa H., Dokya M., Ceramic International, vol. 24, 229, 1998), and amorphous  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}$  (obtained in Example 1,a) were intimately mixed in an agate mortar. Slurry was prepared by adding 10 ml of isopropyl alcohol to the powder mixture (116 mg CGO and 147 mg of amorphous  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}$ ), which was further ultrasonicated in order to reduce the formation of agglomerates. The composite was then heated to 423 K (about 150°C) for solvent evaporation, followed by reduction at 773 K (about 500°C) for 0.5 h under hydrogen flux. The formation of the NiCu alloy on CGO was confirmed by X-ray diffraction (see Figure 5 wherein \* is for  $\text{Ni}_{0.58}\text{Cu}_{0.42}$ , and ° is for CGO).

#### Example 5

##### Preparation of a $\text{Ni}_{0.5}\text{Cu}_{0.5}$ -CGO (50:50 w/w) cermet (reference preparation)

1g of CGO powder was impregnated according to what taught by H. Kim, et al., *supra*, with a 50 ml aqueous solution of  $\text{Cu}(\text{NO}_3)_2 \cdot \text{H}_2\text{O}$  (280 mg) and  $\text{Ni}(\text{NO}_3)_2 \cdot \text{H}_2\text{O}$  (310 mg). Finally the layer was again heated to 500°C and reduced in flowing  $\text{H}_2$  at 900°C.

XRD analysis showed (Figure 6) that the NiCu/CGO cermet prepared according to said method show particle size higher and surface area lower than that obtained in Example 4 (as from comparison with Figure 5). Table 2 sets forth the comparison by numbers.

Table 2  
Particle size (d) and surface area (MSA) of NiCu alloys in NiCu-CGO cermets

Preparation process	d / nm	Surface area (m <sup>2</sup> /g)
Example 4	15	43
Example 5	26	26

#### Example 6

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#### Cell preparation

A cell was fabricated having a CGO electrolyte, a LSM/CGO layer as a cathode and a NiCu-CGO layer of Example 4 as an anode.

10 The CGO electrolyte (~500  $\mu\text{m}$ , >90% theoretical density) was prepared by uniaxial pressing at 300 MPa of a  $\text{Ce}_{0.9}\text{Gd}_{0.1}\text{O}_{1.95}$  powder obtained as in Example 4. Before use for pellet preparation, the powder was thermally treated at 1050°C for 1 h. The pellet was thermally treated at 1550°C for 3 hrs.

15 As the cathode, a 30  $\mu\text{m}$  LSM/CGO layer (50:50%wt) was deposited by a painting process on one side of the pellet and fired at 1250°C for 1 hour in air to assure good bonding to the electrolyte. The slurry used was composed of 100 mg CGO synthesized powder and 100 mg LSM ( $\text{La}_{0.6}\text{Sr}_{0.4}\text{MnO}_3$ , Praxair) both intimately mixed and dispersed in 1.5 ml of isopropanol.

20 A 20  $\mu\text{m}$  anodic cermet layer of amorphous  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}$ -crystalline CGO (50:50%wt) slurry was deposited by painting in one step on the CGO dense layer side of the CGO-LSM/CGO substrate. The slurry was prepared by dispersing 100 mg of amorphous  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}_x$  and 100 mg of synthesized CGO powders in 1.5 ml of isopropylalcohol (carlo Erba). The total amount of deposited metal phase was 2.5 mg/cm<sup>2</sup>. This was dried at 423 K (about 150°C) to remove the solvent.

25 A 5  $\mu\text{m}$  Au (Hereus) film, to be used as the anodic current collector in the electrochemical cell, was then deposited by painting on the anodic layer, and the whole assembly was heated at 150°C for solvent evaporation. Two Au wires on each side were allocated for sampling current and potential.



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Also, a 5  $\mu\text{m}$  thin Pt (Engelard) film, to be used as the cathodic current collector in the electrochemical cell, was then deposited by painting on the cathodic layer. A Pt reference electrode was allocated on the cathodic side to allow operation of the device under half-cell configuration.

5        The cell (0.5  $\text{cm}^2$  active area) was mounted on an alumina tube and sealed with quartz adhesive.

Finally the system was heated at 500°C for 1h in air to allow formation of a crystalline  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}_x$  oxide. Inert gas (He) was passed through the anode before hydrogen supply. An hydrogen stream flow rate (50  $\text{cc min}^{-1}$ )  
10        was fed to the anode at 500°C to assure the alloy formation.

#### Example 7

##### Characterization in half-cell configuration (hydrogen fuelled anode)

Electrochemical evaluation of the performance of a solid oxide fuel cell according to Example 6, fed with hydrogen was carried out.

15        The hydrogen flow rate was 50  $\text{cc min}^{-1}$ , and static air was used as oxidant. No humidification was used for the anode stream.

The cell was conditioned for at least 1 h in hydrogen at 800°C before recording the polarization curves and ac-impedance spectra.

20        Electrochemical experiments were carried out both under galvanostatic and potentiostatic controls by using an AUTOLAB Ecochemie potentiostat/galvanostat and impedance analyser. The polarization data were collected under steady state conditions. Ac-impedance spectra were collected in the range 1 MHz-1 mHz with a 20 mV rms sinusoidal signal under open circuit conditions. A four-electrode configuration was used in all cases. In half-cell  
25        experiments, one potential probe was connected to a non-polarized reference electrode and the overpotential of the working electrode was measured against this reference.

30        Raw half-cell data of the  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}/\text{CGO}$  cermet anode/CGO electrolyte interface indicated that this anode is active for the electrochemical oxidation of dry  $\text{H}_2$ , as depicted in Figure 7.

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IR-free data (Figure 8) show very low overpotentials, less than 50 mV for  $j = 500 \text{ mA cm}^{-2}$ . The curves recorded in presence of hydrogen show no activation (kinetic) control. These data indicate that the  $\text{Ni}_{0.58}\text{Cu}_{0.42}/\text{CGO}$  cermet anode combined with thin CGO electrolyte achieves high performances for the oxidation of dry hydrogen.

#### Example 8

##### Characterization in half-cell configuration (dry methane fuelled anode)

Electrochemical evaluation of the performance of a solid oxide fuel cell according to Example 6, fed with hydrogen was carried out.

Methane flow rate was  $50 \text{ cc min}^{-1}$ , and static air was used as oxidant. No humidification was used for the anode stream.

The cell was conditioned for at least 1 h in methane at  $800^\circ\text{C}$  before recording the polarization curves and ac-impedance spectra.

Electrochemical experiments were carried out as in Example 8, but testing the cell at three different temperatures, i.e.  $800^\circ\text{C}$ ,  $700^\circ\text{C}$  and  $600^\circ\text{C}$ , sequentially.

Raw half-cell data of the  $\text{Ni}_{0.58}\text{Cu}_{0.42}\text{O}/\text{CGO}$  cermet anode/CGO electrolyte interface indicated that this anode is active for the electrochemical oxidation of dry methane, as depicted in Figure 7.

IR-free data (8) show for  $j = 500 \text{ mA cm}^{-2}$  overpotentials of 250 and 350 mV vs. the reversible potential for  $\text{H}_2$  oxidation at  $800^\circ\text{C}$  and  $700^\circ\text{C}$  respectively. The curves recorded in presence of methane show a slight activation (kinetic) control that is not observed in presence of hydrogen. These data indicate that the  $\text{Ni}_{0.58}\text{Cu}_{0.42}/\text{CGO}$  cermet anode combined with thin CGO electrolyte achieve high performances for the oxidation of dry methane.

#### Example 9

##### Analysis of morphology and carbon deposition characteristics of the anodic cermet

The overall process of formation of the  $\text{Ni}_{0.58}\text{Cu}_{0.42}$  alloy/CGO cermet in a cell prepared according to Example 6, and the modifications occurring after its exposure to dry methane, according to Example 8 under electrochemical

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The cermet anode of the present invention is able to make the solid oxide fuel cell comprising it to operate with a wide selection of fuels. Especially said cermet, thanks to its characteristics of ionic and electronic conductivity, and surface area and catalytic activity, permits the use of dry methane as fuel

5 for electrochemical apparatus, without any sign of carbon deposition.

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# CLAIMS

1. Solid oxide fuel cell including
  - a cathode;
  - at least an electrolyte membrane, and
  - 5 - an anode comprising a ceramic material and a Cu/Ni alloy, said Cu/Ni alloy having an average particle size not higher than 20 nm.
2. Solid oxide fuel cell according to claim 1 wherein said Cu/Ni alloy has an average particle size not higher than 16 nm.
3. Solid oxide fuel cell according to claim 1 wherein said Cu/Ni alloy has a  
10 mean surface area higher than 20 m<sup>2</sup>/g.
4. Solid oxide fuel cell according to claim 3 wherein said Cu/Ni alloy has, a mean surface area higher than 30 m<sup>2</sup>/g.
5. Solid oxide fuel cell according to claim 4 wherein said Cu/Ni alloy has, a mean surface area higher than 40 m<sup>2</sup>/g.
- 15 6. Solid oxide fuel cell according to claim 1 wherein said Cu/Ni alloy has a copper content of from 1% by weight to 99% by weight.
7. Solid oxide fuel cell according to claim 6 wherein said Cu/Ni alloy has a copper content of from 30% by weight to 70% by weight.
8. Solid oxide fuel cell according to claim 7 wherein said Cu/Ni alloy has a  
20 copper content of from 40% by weight to 60% by weight .
9. Solid oxide fuel cell according to claim 1 wherein said Cu/Ni alloy has a nickel content of from 1% by weight to 99% by weight.
10. Solid oxide fuel cell according to claim 6 wherein said Cu/Ni alloy has a nickel content of from 30% by weight to 70% by weight.
- 25 11. Solid oxide fuel cell according to claim 7 wherein said Cu/Ni alloy has a nickel content of from 40% by weight to 60% by weight .
12. Solid oxide fuel cell according to claim 1 wherein said Cu/Ni alloy comprises an additional metal.
13. Solid oxide fuel cell according to claim 1 wherein said ceramic material is  
30 selected from yttria-stabilized zirconia (YSZ), cerium gadolinium oxide (CGO), samarium-doped ceria (SDC), mixed lanthanum and gallium oxides.

14. Solid oxide fuel cell according to claim 1 wherein said ceramic material has a particle size not higher than 50 nm.
15. Solid oxide fuel cell according to claim 1 wherein said ceramic material has a particle size from 1 nm to 25 nm.
- 5 16. Solid oxide fuel cell according to claim 1 wherein said ceramic material is doped with at least one cation selected from calcium, magnesium, strontium, lanthanum, yttrium, ytterbium, neodymium and dysprosium.
17. Solid oxide fuel cell according to claim 13 wherein said ceramic material is cerium gadolinium oxide (CGO).
- 10 18. Solid oxide fuel cell according to claim 1 performing in substantially dry hydrocarbon.
19. Cermet comprising a ceramic material and a Cu/Ni alloy having a particle size not higher than 20 nm.
20. Process for preparing a cermet including a ceramic material and a  
15 metallic material comprising a Cu/Ni alloy, said process comprising the steps of:
- a) producing a precursor of the metallic material;
  - b) producing the ceramic material;
  - c) combining said precursor and ceramic material to obtain a composite
  - 20 d) reducing said composite
- wherein step a) comprises the phases of
- a-1) dissolving a hydrosoluble salt of Ni and a hydrosoluble salt of Cu in water;
  - a-2) adding a chelating agent to the solution resulting from step
  - 25 a-1);
  - a-3) adding an oxidizing agent to the solution resulting from step a-2);
  - a-4) isolating said precursor.
21. Process according to claim 20 wherein step b) comprises the phases  
30 analogous to those from a-1) to a-4).

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22. Process according to claim 20 comprising the phases of adjusting the pH of the solution resulting from phase a-2) at a value higher than about 5.

23. Process according to claim 20 wherein phase d) is carried out with hydrogen at a temperature ranging between about 400°C and about 1000°C.

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# ABSTRACT

A solid oxide fuel cell including a cathode, at least an electrolyte membrane, and an anode comprising a ceramic material and a Cu/Ni alloy, said Cu/Ni alloy having an average particle size not higher than 20 nm.

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(Fig.1)

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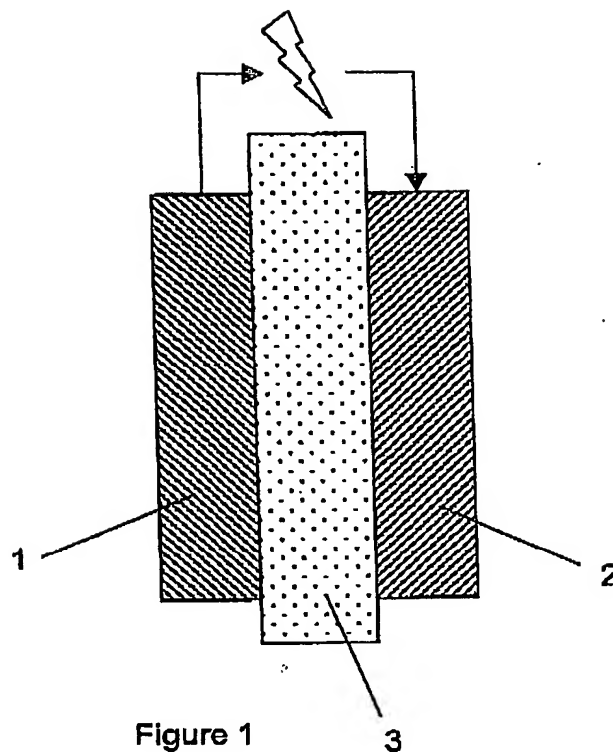


Figure 1

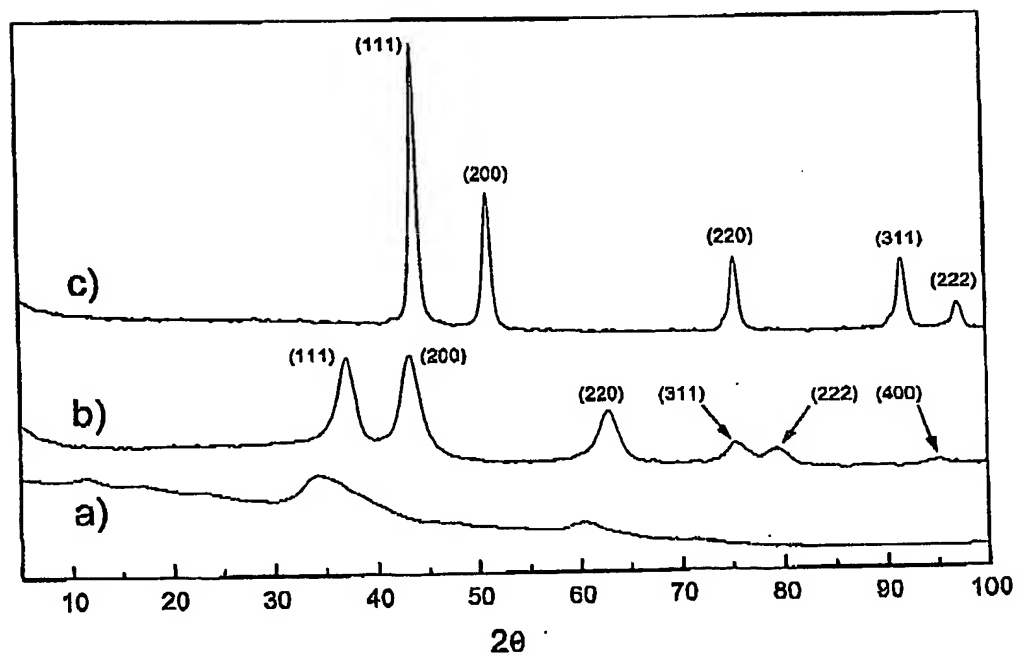


Figure 2

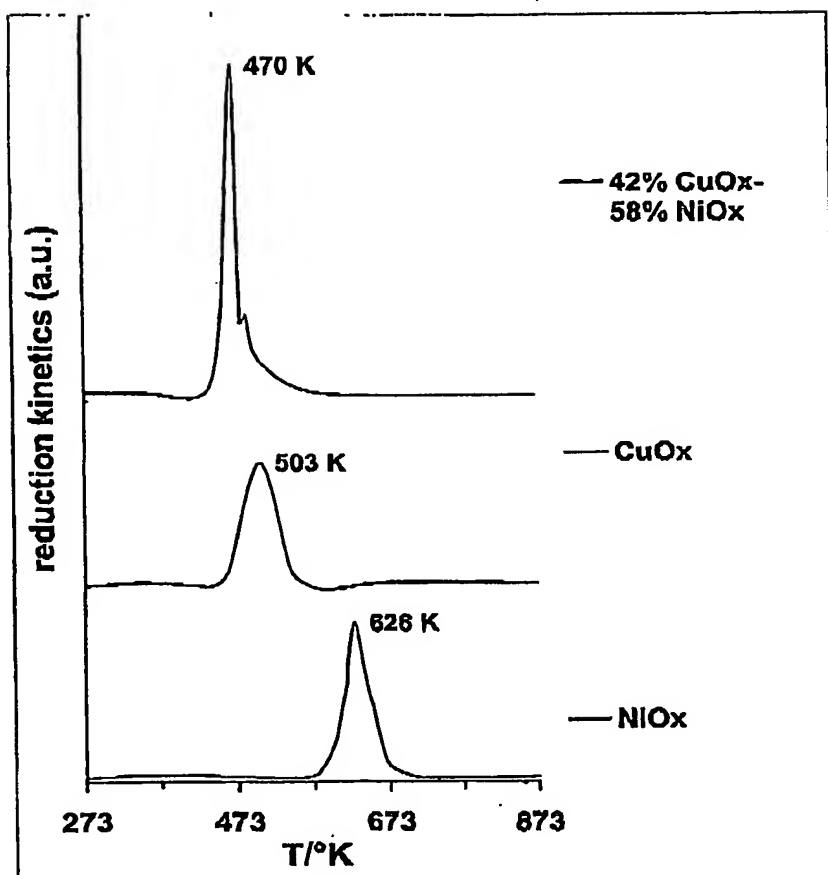
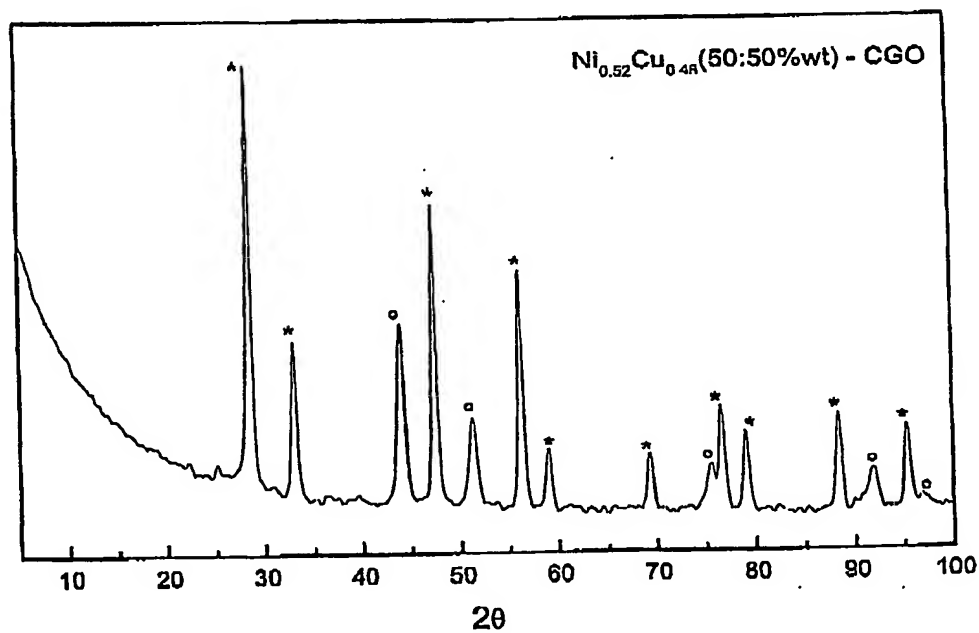
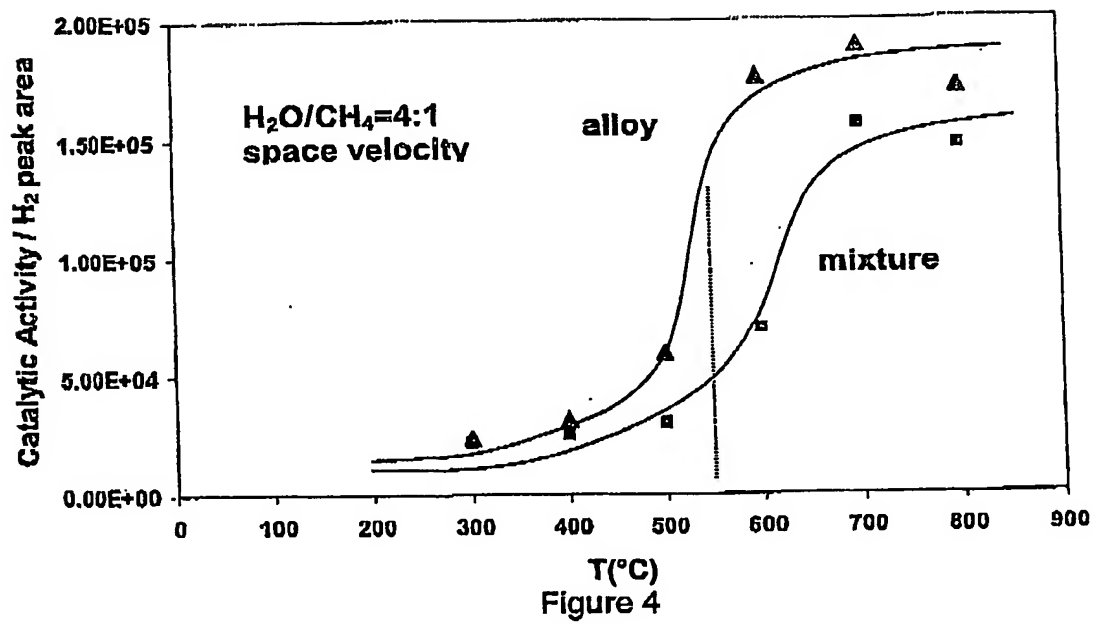


Figure 3



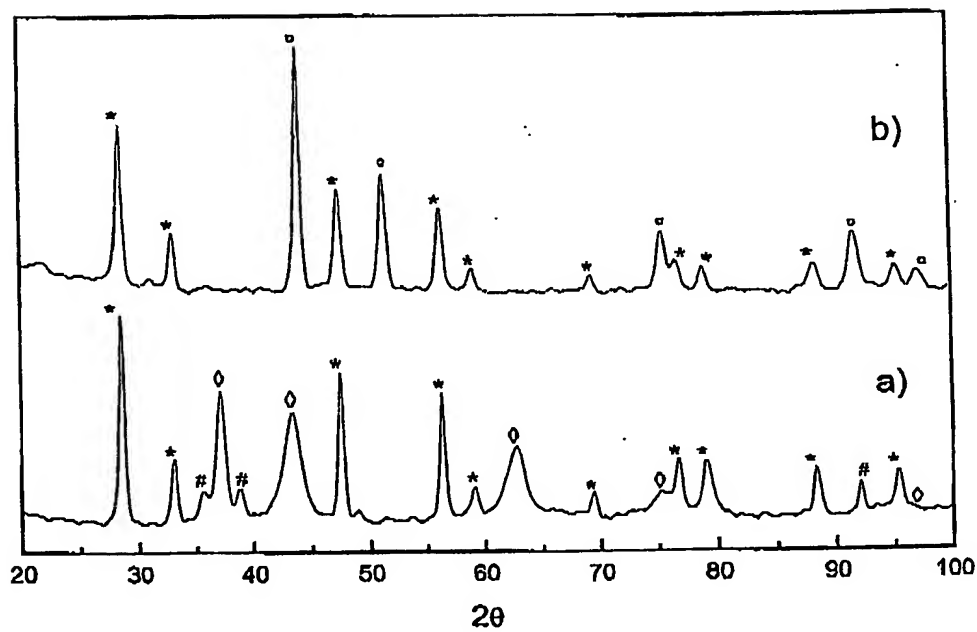


Figure 6

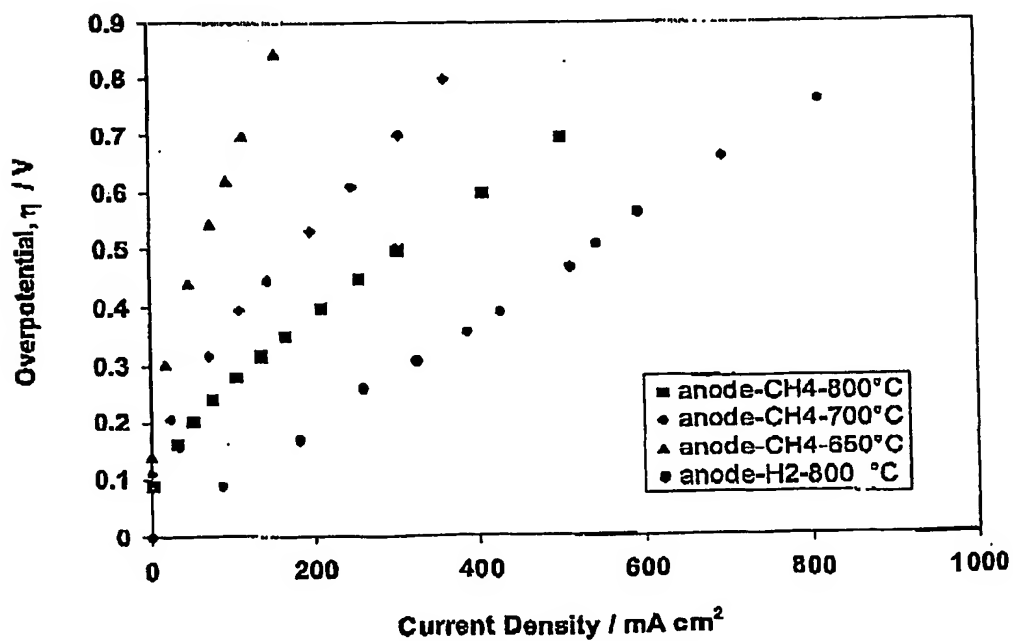


Figure 7

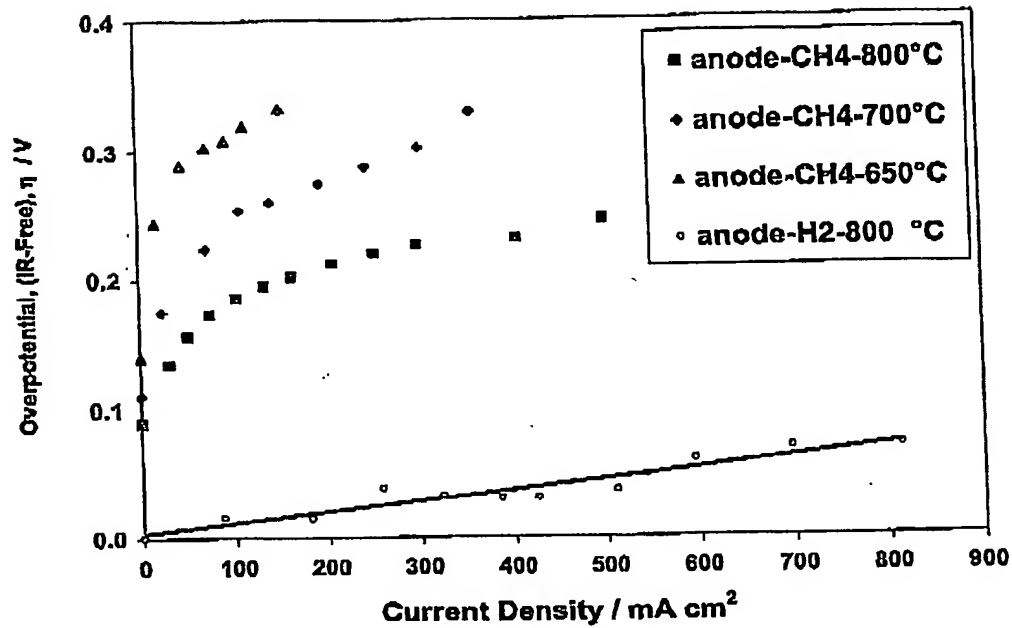


Figure 8

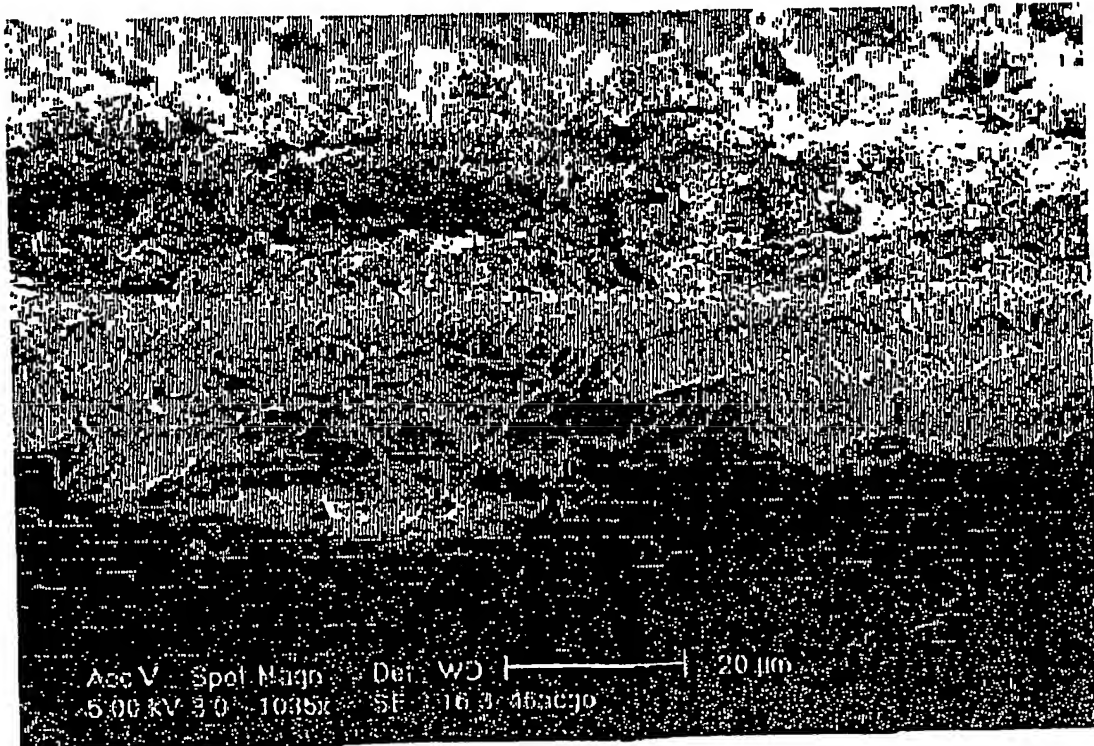


Figure 9

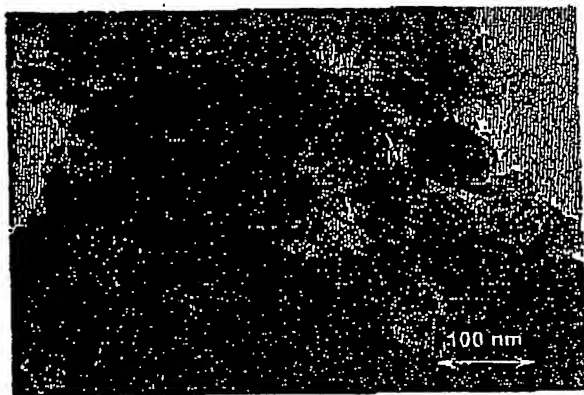


Figure 10a

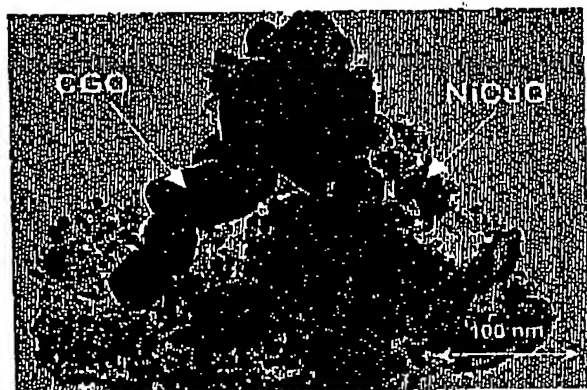
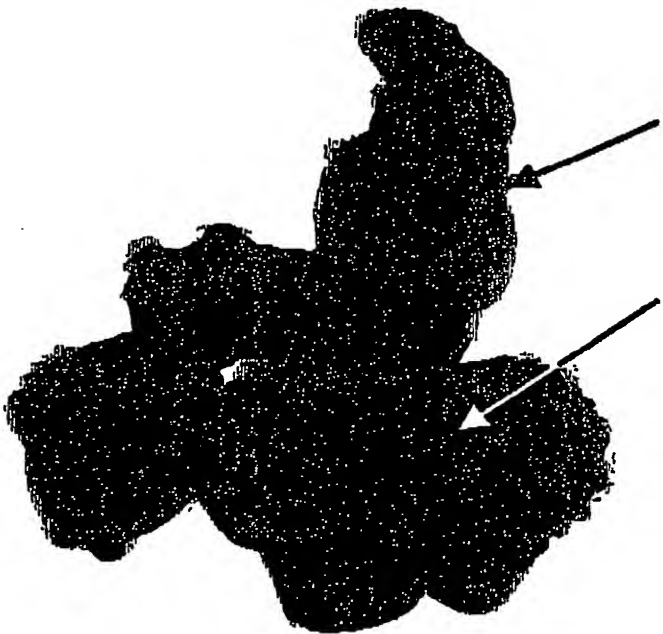


Figure 10b

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CGO particle

The image shows a dark, irregularly shaped cluster of particles. Two arrows point to specific regions within this cluster. The upper arrow points to a more elongated, vertical section, while the lower arrow points to a more horizontal, rounded section.

NiCu particle

Figure 11

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